

Problema0232: Calcula a composición centesimal do ácido sulfúrico, H_2SO_4

$$M_m(H_2SO_4) = 2 \cdot 1\text{ g} + 32\text{ g} + 4 \cdot 16\text{ g} = 98\text{ g}$$

$$\%H = \frac{2\text{ g}}{98\text{ g}} \cdot 100 = \underline{2,04\%H}$$

$$\%S = \frac{32\text{ g}}{98\text{ g}} \cdot 100 = \underline{32,65\%S}$$

$$\%O = \frac{64\text{ g}}{98\text{ g}} \cdot 100 = \underline{65,31\%O}$$

Comproba que a suma das porcentaxes debe dar 100%

$$Suma = 2,04 + 32,65 + 65,31 = 100\%$$