

Problema0233: Calcula a composición centesimal do carbonato de amonio, $(\text{NH}_4)_2\text{CO}_3$

$$M_m[(\text{NH}_4)_2\text{CO}_3] = 2 \cdot 14 \text{ g} + 8 \cdot 1 \text{ g} + 12 \text{ g} + 3 \cdot 16 \text{ g} = 96 \text{ g}$$

$$\%N = \frac{28 \text{ g}}{96 \text{ g}} \cdot 100 = \underline{29,17\%N}$$

$$\%H = \frac{8 \text{ g}}{96 \text{ g}} \cdot 100 = \underline{8,33\%H}$$

$$\%C = \frac{12 \text{ g}}{96 \text{ g}} \cdot 100 = \underline{12,50\%C}$$

$$\%O = \frac{48 \text{ g}}{96 \text{ g}} \cdot 100 = \underline{50,00\%O}$$

Comproba que a suma das porcentaxes debe dar 100%

$$\text{Suma} = 29,17 + 8,33 + 12,50 + 50,00 = 100 \%$$