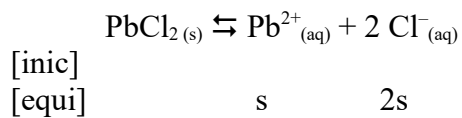


EXEMPLO 12: O  $K_s$  do  $\text{PbCl}_2$  é  $1,7 \cdot 10^{-5}$ . Cal é a súa solubilidade en g/l?



$$K_s = [\text{Pb}^{2+}_{(aq)}] \cdot [\text{Cl}^{-}_{(aq)}]^2 = s \cdot (2s)^2 = 4s^3 = 1,7 \cdot 10^{-5}$$

$$s = \sqrt[3]{\frac{1,7 \cdot 10^{-5}}{4}} = 0,0162 \text{ M} = \frac{0,0162 \text{ mol}}{1 \text{ L}} \cdot \frac{278,1 \text{ g}}{1 \text{ mol}} = \underline{4,51 \text{ g/L}}$$