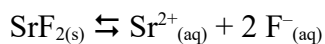
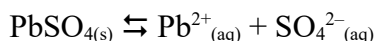


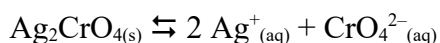
Problema651: Escribe o equilibrio de disociación destes sales, escribe a expresión do produto de solubilidade (K_s), e busca nas táboas de química o valor da constante: SrF_2 , PbSO_4 , Ag_2CrO_4 , NiCO_3 , $\text{Cr}(\text{OH})_3$, Bi_2S_3 , $\text{Pb}_3(\text{PO}_4)_2$.



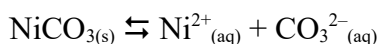
$$K_s = [\text{Sr}^{2+}_{(aq)}] \cdot [\text{F}^{-}_{(aq)}]^2 = 7,9 \cdot 10^{-10}$$



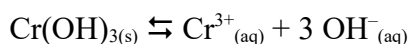
$$K_s = [\text{Pb}^{2+}_{(aq)}] \cdot [\text{SO}_4^{2-}_{(aq)}] = 1,3 \cdot 10^{-8}$$



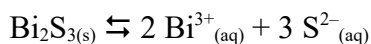
$$K_s = [\text{Ag}^{+}_{(aq)}]^2 \cdot [\text{CrO}_4^{2-}_{(aq)}] = 1,9 \cdot 10^{-12}$$



$$K_s = [\text{Ni}^{2+}_{(aq)}] \cdot [\text{CO}_3^{2-}_{(aq)}] = 1,4 \cdot 10^{-7}$$



$$K_s = [\text{Cr}^{3+}_{(aq)}] \cdot [\text{OH}^{-}_{(aq)}]^3 = 6,7 \cdot 10^{-31}$$



$$K_s = [\text{Bi}^{3+}_{(aq)}]^2 \cdot [\text{S}^{2-}_{(aq)}]^3 = 1,6 \cdot 10^{-72}$$



$$K_s = [\text{Pb}^{2+}_{(aq)}]^3 \cdot [\text{PO}_4^{3-}_{(aq)}]^2 = 1 \cdot 10^{-54}$$