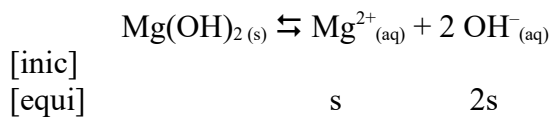


Problema654: Cal é a solubilidade do hidróxido de magnesio en auga pura? $K_s[\text{Mg}(\text{OH})_2] = 8,9 \cdot 10^{-12}$



$$K_s = [\text{Mg}^{2+}_{(aq)}] \cdot [\text{OH}^{-}_{(aq)}]^2 = 8,9 \cdot 10^{-12}$$

$$K_s = [\text{Mg}^{2+}_{(aq)}] \cdot [\text{OH}^{-}_{(aq)}]^2 = s \cdot (2s)^2 = 4s^3 = 8,9 \cdot 10^{-12}$$

$$s = \sqrt[3]{\frac{8,9 \cdot 10^{-12}}{4}} = \underline{1,31 \cdot 10^{-4} M}$$