EXEMPLO 2: Determina o número de oxidación dos elementos que forman as seguintes substancias:

$\overset{+1}{H_2}\overset{-2}{O}$, $\overset{+1}{H}\overset{-1}{Cl}$, $\overset{+1}{Na}\overset{-1}{Cl}$, $\overset{+2}{Zn}\overset{-1}{Cl_2}$, $\overset{+2}{Zn}\overset{-2}{O}$, $\overset{+1}{H_2}\overset{+6}{S}\overset{-2}{O_4}$, $\overset{+6}{S}\overset{-2}{O_4^{2-}}$, $\overset{+1}{H}\overset{+5}{N}\overset{-2}{O_3}$, $\overset{+5}{N}\overset{-2}{O_3^{-}}$, $\overset{+1}{H_2}\overset{-2}{S}$, $\overset{-2}{S^{2-}}$,

$\overset{+1}{K}\overset{+7}{Mn}\overset{-2}{O_4}$, $\overset{+4}{Mn}\overset{-2}{O_2}$, $\overset{+5}{V}\overset{-2}{O^{3+}}$, $\overset{+5}{P_4}\overset{-2}{O_{10}}$, $\overset{+1}{Li}\overset{-1}{H}$, $\overset{+5}{I}\overset{-1}{Cl_5}$, $\overset{+3}{N}\overset{-1}{F_3}$, $\overset{+4}{S}\overset{-2}{O_3^{2-}}$, $\overset{+3}{S_2}\overset{-2}{O_4^{2-}}$, $\overset{+6}{Cr_2}\overset{-2}{O_7^{2-}}$.